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Biological Effect of Cobalt Oxide Nanoparticles Synthesized from (*Cinnamon* and *Nigella sativa*) Extracts and Chemical Method: A Comparative Study

Ghasaq Z. Alwan[1,*], Emtinan Mohammed Jawad[1,2], Hana H. Kareem[1,3] & Raed Jumah Almansouri[4]

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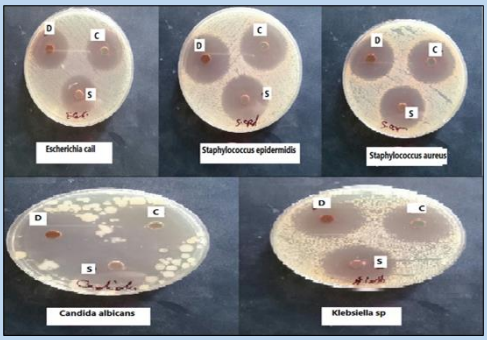
Abstract: This study involved the synthesis of cobalt oxide (Co_3O_4) nanoparticles (NPs). Using a sustainable green utilization of *Cinnamon* (*Cinnamomum zeylanicum*) and *Nigella sativa* extracts, alongside a conventional chemical synthesis method for comparison. Comprehensive structural, morphological, and optical characterizations confirmed the successful synthesis of Co_3O_4 NPs with distinct properties. XRD analysis revealed smaller crystalline sizes for green-synthesized NPs (5.49 nm for *Cinnamon* and 7.22 nm for *Nigella sativa*) compared to chemically synthesized ones (10.6 nm), while FTIR identified functional groups contributing to enhanced chemical and biological activity. FESEM and AFM analyses showed improved surface topography, reduced roughness, and higher effective surface areas in biogenically NPs. UV-visible spectroscopy demonstrated enhanced electronic properties, with band gaps of 1.61 eV and 1.59 eV for *Cinnamon* and *Nigella sativa*-derived NPs, respectively, as compared to 1.43 eV for chemically synthesized NPs. Biological assays revealed superior antimicrobial activity of green-synthesized NPs, particularly *Cinnamon*-derived NPs against *Staphylococcus aureus* (34 mm), *Escherichia coli* (30 mm), and *Nigella sativa*-derived NPs against *Candida albicans* (41 mm). The synergistic action between plant-derived compounds such as Cinnamaldehyde and Thymoquinone and cobalt oxide nanoparticles enhanced their antimicrobial efficacy. In conclusion, the use of plant extracts for synthesizing Co_3O_4 NPs offers a sustainable, economical substitute for the chemical approach, producing nanoparticles.



Keywords: Antibacterial activity, *Cinnamon*, Co_3O_4 Nps, green synthesis, Plant extract.

Introduction

The creation and utilization of metal and metal oxide nanoparticles have attracted considerable interest because of their unique characteristics, which result from their structural features, size, shape, and overall dimensions [1]. These nanoparticles have intrigued researchers for over a century, playing a crucial role across numerous scientific and industrial domains. Transition metal oxides at the nanoscale are particularly noteworthy as a class of materials with diverse applications, including their roles as catalysts, electrodes, semiconductors, sensors, components in solar cells, and ceramic materials, among others [2-4].

The proliferation of antibiotic-resistant bacterial strains exacerbates a major risk to public health by the widespread appearance of infectious illnesses. Antibiotics have been the main line of defense against infections for many years in both community and *medical* settings [5, 6].However, advancements in nanobiotechnology, especially the tailored synthesis of metal oxide nanomaterials with specific morphologies, exhibit potential for the evolution of new antibacterial substances. Nanoparticles, in particular, have generated significant attention due to their functional characteristics, which are heavily impacted due to their dimensions and large surface area relative to volume [7, 8]. The utilization of inorganic materials offers distinct advantages, including chemical stability, safety, and biocompatibility [9]. A growing body of literature highlights the antibacterial efficacy of metal oxide nanoparticles, emphasizing their high potential in combating bacterial infections [10] and the chain reactions [11]. The potency of antioxidants is strongly influenced by their concentration and the chemical states of the antioxidants with which they interact [12].Various antioxidants have been identified as essential for scavenging toxic free radicals in biological systems, mitigating the oxidative stress generated during these processes [13, 14]. Due to their high surface-to-volume ratio, nanostructured materials are anticipated to outperform bulk materials as free radical scavengers, offering enhanced protective capabilities. Nanoparticles' antimicrobial properties also hinge on the structural characteristics of the target organisms. Gram-positive bacteria are typically more resistant to antibacterial treatments than Gram-negative bacteria because of their thicker cell walls. Nanoparticles have

1 Department of Physics, College of Science, Mustansiriyah University, Baghdad, Iraq.
* Corresponding author email: ghasaqz.alwan@uomustansiriyah.edu.iq
2 E-mail: imtenanmj@uomustansiriyah.edu.iq
3 E-mail: hanakr2007@uomustansiriyah.edu.iq
4 Directorate of education of Al-Rasafa, Ministry of Education, Baghdad, Iraq. raedpcms@gmail.com

demonstrated encouraging antibacterial properties despite this. When nanoparticles interact with bacterial cells, they damage the integrity of the membrane, resulting in holes and pits that weaken the cell's structure. This interference suppresses respiratory chain enzymes, prevents DNA replication, and eventually results in cell death [15-17]. Several chemical agents, like hydrazine, sodium borohydride, aldehydes, and citrates, as well as radiation sources and functional polymers that operate as reducing agents, are typically used in the traditional synthesis of nanoparticles. Nevertheless, these approaches have a large impact on the environment, are expensive, and use a lot of energy and resources [18, 19]. Recent studies have turned to plant extracts as a sustainable and eco-friendly substitute for nanoparticle manufacturing in response to these issues. In addition to being less harmful, plant-based techniques are safer for human use [2, 20]. Numerous biochemicals found in botanical extracts, including proteins, amino acids, terpenoids, alkaloids, phenolics, and aldehydes, stabilize and enhance the conversion of metal ions into nanostructures. These extracts have negligible contamination risks, are easily handled, and are suitable for industrial synthesis scaling up. Numerous studies have revealed the successful formation of nanoscale metal particles using a range of plant extracts [21, 22]. Achieving exact control over the nanoparticles' size and structure through reaction condition tweaking is a crucial difficulty in their production. Metabolites in plant extracts also serve as capping agents during synthesis, preventing nanoparticle aggregation [23-25]. Since various plant extracts contain multiple amounts and configurations of chemical substances, the features of the created nanoparticles are set by the plant sources [26]. For the manufacturing of CoO_4 NPs, *cinnamon* (*Cinnamomum zeylanicum*) was used as the plant source in this investigation. *Cinnamon* is widely used as a seasoning agent and in herbal remedies to treat digestive ailments such as diarrhea. Its extract contains a high concentration of bioactive molecules such as cinnamaldehyde, eugenol, and linalool, which exhibit potent antioxidant activity [27]. Additionally, *Nigella sativa*, an herbaceous seasonal plant from the Ranunculaceae family, was also employed for Co_3O_4 NP biosynthesis [28].Cobalt oxides, comprising cobaltous oxide (CoO), cobaltic oxide (Co_2O_3), and cobaltosic oxide (Co_3O_4), possess unique physicochemical properties that support a wide range of industrial applications [29]. Among these, Co_3O_4 has achieved particular care due to its distinctive morphology and redox properties, as well as its nonstoichiometric polymorphic nature [30]. Several studies have investigated biogenic synthesis using plant extracts of cobalt oxide nanoparticles. Dubey *et al. [31]* synthesized cobalt oxide NPs using latex from *Calotropis procera*, while Bibi *et al. [32]* employed *Punica granatum* peel extract for cobalt oxide fabrication. Similarly, Akhlaghi *et al.* [33] used *Trigonella foenum-graecum (fenugreek)* leaf extract, and Saeed *et al. [34]* utilized root extracts from *Ziziphus oxyphylla Edgew*.. Alaghemand *et al. [28]* fabricated silver nanoparticles by *Nigella sativa* L. extract, while Anjum *et al. [35]* presented the findings on formation of cobalt oxide NPs using *Cinnamomum zeylanicum*. The purpose of this study is to develop a rapid and efficient method for synthesizing Co_3O_4 nanoparticles using plant extracts from *Cinnamomum zeylanicum* and *Nigella sativa* L., alongside a comparative chemical synthesis approach without plant extracts. The study further evaluates the inhibition of these NPs on bacterial and fungal pathogens.

Materials and Methods

Materials

Cobalt nitrate $Co(NO_3)_2$, M.W. 182.92 g/mol was obtained from Alpha Chemika (India). Sodium hydroxide (NaOH, M.W. 40 g/mol), as well as *cinnamon* bark and *Nigella sativa* seeds, were procured from local markets.

Preparation of Co_3O4 NPs with Nigella sativa extract

Two grams of *Nigella sativa* seeds were ground using an electric grinder for ten minutes and sieved through a 58-micron mesh. The synthesized powder was mixed into 100 mL of deionized water and then stirred on a magnetic mixer at 50°C for 30 minutes, yielding a light yellow, transparent solution. Upon cooling to room temperature, the extract was strained four times using Whatman filter paper no. 0.5 separately, 9.15 g of cobalt nitrate was dissolved in 100 mL of deionized water (0.5 M) and stirred on a magnetic mixer at 50°C till a transparent dark red solution was obtained. Then, 35 mL of the *Nigella sativa* extract was poured in drops into the cobalt nitrate solution, being stirred continuously, culminating in the creation of a reddish-brown aqueous solution. As shown in Fig. (1-A).

Preparation of Co_3O_4 NPs with Cinnamon extract

Two grams of *cinnamon* were ground using an electric grinder for ten minutes, next sieved via a 58-micron mesh. The formed powder was mixed into 100 mL of deionized water and then stirred on a magnetic mixer at a temperature of 50°C for 30 minutes, producing a yellow solution. After reaching room temperature, the extract was strained four times using Whatman filter paper no. 0.5 separately, 9.15 g of cobalt nitrate was dissolved in 100 mL of deionized water (0.5 M) and stirred on a magnetic mixer at 50°C till a clear ruby-red solution was obtained. Thereafter, 35 mL of the plant extract was poured in drops into the cobalt nitrate solution, being stirred continuously. The mixture was distilled until the solution's color changed to reddish-brown. As illustrated in Fig. (1-B).

Preparation of Co_3O_4 NPs by chemical method without using any plant extracts

A total of 9.1 g of cobalt nitrate ($Co(NO_3)_2$) was mixed in 100 mL of deionized water and stirred on a magnetic mixer at 50°C for 30 minutes, producing a red solution upon complete dissolution. In a separate preparation, 4 g of sodium hydroxide (NaOH) was mixed in 100 mL of deionized water to create a 1 M solution and stirred magnetically at 50°C until fully dissolved. The NaOH solution was then added gradually and dropwise to the cobalt nitrate solution at a steady rate. A distinct color change was observed, with the solution stabilizing to a blackish-blue hue after the addition of 20 mL of the NaOH solution. As shown in Fig.(1-C).

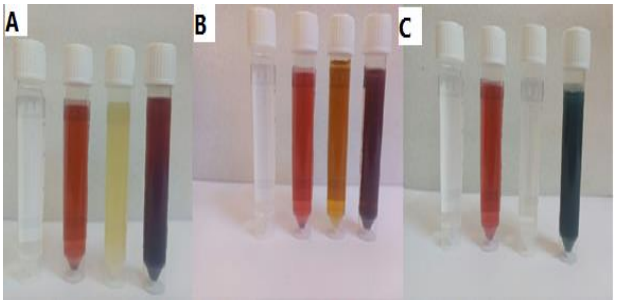


Figure (1): Co_3O_4 NPs formation steps :(A) using *Nigella sativa* extract. (B) using *Cinnamon* extract (C) using chemical method without using any plant extracts.

Antimicrobial Activity Measurement

The antimicrobial effect of cobalt oxide nanoparticles (CO_3O_4 NPs) synthesized chemically, combined with *cinnamon* and *Nigella sativa* extracts, was evaluated under aerobic conditions with efficient agar diffusion in well plates against various pathogens. The effectiveness of the samples in inhibiting bacterial growth, fungi, and yeast was tested against *Staphylococcus aureus*, Staphylococcus epidermis, *Escherichia*

coli, *Klebsiella sp*., and *Candida albicans*. Nutrient agar plates were prepared and inoculated with the test organisms, which were evenly spread using a glass spreader. Agar wells measuring 6 mm in diameter were created in the Muller–Hinton substrate, and 100 µL of each test sample, including Co_3O_4 NPs, *cinnamon* extract, and *Nigella sativa* extract, were introduced. The plates were incubated under optimal conditions, with fungal and bacterial samples that were cultured for 72 hours at 30 °C and for 24 hours at 37 °C, respectively. The range of the zone of inhibition (ZOI) was calculated with a millimeter scale to determine the antibacterial activity. This technique allowed the diffusion of the test substances into the agar, effectively inhibiting the growth of the test organisms. The experiment was conducted in the Life Sciences Department Laboratory at Mustansiriyah University, demonstrating the potential of these materials as antimicrobial agents.

Characterization

The characterization of the as-prepared Co_3O_4 nanoparticles (NPs) was conducted using various analytical techniques. UV–visible spectrophotometry was performed with a DU-8800D spectrophotometer (China) within the wavelength range of 190–1100 nm. FTIR spectra were recorded using an attenuated total reflection Fourier-transform infrared spectrophotometer (ATR-FTIR, ALPHA - BRUKER). The morphological properties of the Co_3O_4 NPs were examined using Field Emission Scanning Electron Microscopy (FESEM, MIRA3 TESCAN). X-ray diffraction (XRD) patterns were recorded with a SHIMADZU XRD-6000 instrument (Japan) employing CuKα radiation (λ = 1.5406 Å). Atomic Force Microscopy (AFM) images were acquired using a Digital Instruments system (Nanoscope III and Dimension 3100).

Results and Discussion

XRD Analysis

As seen in Fig. 2, the XRD patterns for Co_3O_4 nanoparticles (NPs) were prepared using *Cinnamon* extract, *Nigella sativa* extract, and the chemical method without plant extracts. Several distinct peaks were observed in all patterns for Co_3O_4 at 2θ = 18.9°, 29.5°, 31.5°, 36.9°, 44.7°, 59.3°, and 65.2°, corresponding to the (111), (210), (220), (311), (400), (511), and (440) planes, respectively. These peaks are characteristic of the spinel-type cubic phase of Co_3O_4 (JCPDS Card No. 01-080-1535) [36]. The Co_3O_4 NPs prepared with plant extracts exhibited similar peaks, with slight shifts due to lattice strains. Interestingly, a peak at 2θ = 15.1°, corresponding to the (101) plane, appeared only in the chemically synthesized Co_3O_4 NPs [37, 38]. This absence in the plant-assisted samples is attributed to the interaction of organic compounds and other materials present in plant extracts with cobalt during synthesis, which impacts the atomic structure, resulting in a partially amorphous phase [38, 39]. Additionally, the small size of the plant-assisted NPs may contribute to peak broadening in the XRD patterns, causing certain peaks to weaken or disappear at specific angles [37]. The intensity of the peaks for plant-assisted NPs was lower compared to those prepared by the chemical method, indicating reduced crystallinity. The crystallite size of Co_3O_4 NPs was computed using the Debye-Scherrer formula, $D=0.94\lambda/\beta\cos\theta$, where λ is the wavelength of x-ray (1.5406 A) for CuKα radiation, β is the full width at half maximum and Θ is the peak position [40]. The average crystallite size was determined to be 5.49 nm for *Cinnamon* extract-assisted NPs, 7.225 nm for *Nigella sativa* extract-assisted NPs, and 10.6 nm for chemically synthesized NPs [41].Table1summarizes XRD characterization and computation of Various Parameters for biogenic Co_3O_4 NPs.

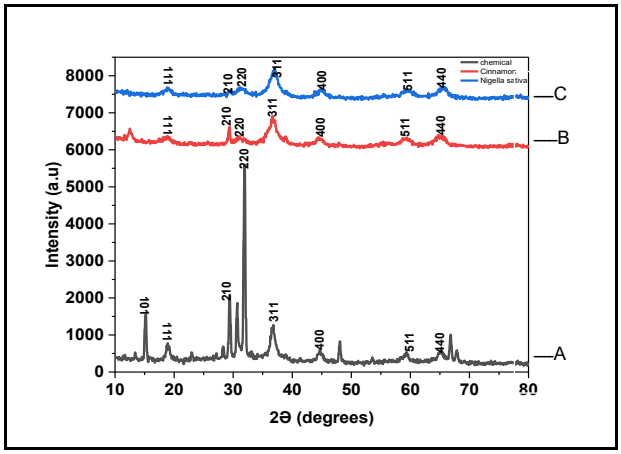


Figure (2): X-ray diffraction patterns of Co_3O_4 NPs with by (A) chemical method, (B) *Cinnamon* extracts, (c) *Nigella sativa* extract.

Table (1): XRD Analysis and calculation of various parameters for Co_3O_4 NPs using chemical method

Sr No.	2θ (Deg.)	FWHM (Deg.)	FWHM (radians.)	dhkl.(Å)	hkl	D (nm)	For unit cell edges: a = b = c (Å)
1	31.8000	0.5000	0.0085	2.8117	220	16.5	
2	36.8440	1.7000	0.0289	2.4376	311	4.9	
3	29.3700	0.5500	0.00935	3.0386	210	14.9	
4	44.6000	1.4000	0.02436	2.0300	400	6.1	
5	15.25	0.62	0.0154	5.8243	101	12.9	7.952
6	18.92	0.68	0.01156	4.6867	111	11.8	
7	59.37	1.48	0.0256	1.5554	511	6.2	
8	65.2	1.77	0.03009	2.0300	400	4.9	

Analysis using FT-IR

The FT-IR spectral range was maintained between 400 and 4000 cm^{-1}. The FT-IR spectra of the biosynthesized Co_3O_4 nanoparticles. Shown in Fig. 3, provides insight into the various vibrational modes present. A prominent peak observed at 3431 cm^{-1} is attributed to O–H stretching [42]. Additionally, a slightly elevated peak at 2066 cm^{-1} corresponds with C–H stretching vibrations within the nanoparticles. The sharp and intense peak at 1638 cm^{-1} is attributed to H–O–H bending vibrations in water molecules, which also indicates the presence of moisture. Furthermore, the δN-H (amide II) group is confirmed by a distinct peak at 1380 cm^{-1} in the biosynthesized Co_3O_4 NPs [41]. Finally, the peak at 661 cm^{-1}associated with the tetrahedral interaction of Co^{2+} with O^{2-} [43].

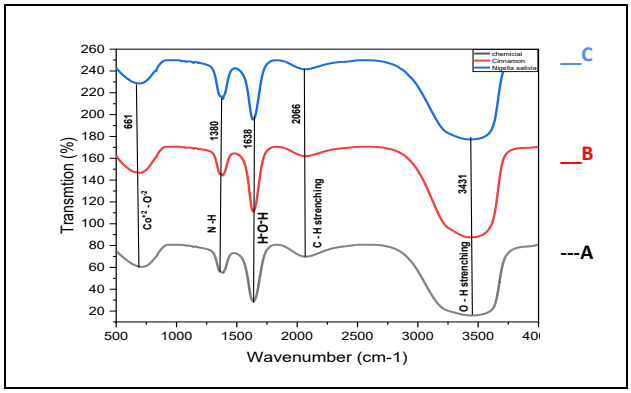


Figure (3): FTIR spectrum of Co_3O_4 NPs with (A) Chemical method, (B) *Cinnamon* extracts (C) *Nigella sativa* extract.

Ultraviolet-visible absorption spectra

The Ultraviolet-visible absorption spectra and the energy gap of Co_3O_4 nanoparticles (NPs) synthesized with *Nigella sativa* extract, *Cinnamon* extract, and the chemical method are shown in Fig. 4. The energy gaps were calculated for the different synthesis methods using the Tauc formula [44], providing valuable insights into their optical properties. For Co_3O_4 NPs made using *Nigella sativa* extract, the energy gap was found to be 1.59 eV, but the gap for the *cinnamon* extract was somewhat greater at 1.61 eV. Conversely, the Co_3O_4 NPs that were chemically produced displayed a smaller energy gap of 1.43 eV [45]. These discrepancies highlight how the materials' electronic structures differ. Additionally, all three approaches yielded strong absorption edges for Co_3O_4 NPs, which were compatible with the XRD data showing that the quantum confinement effect caused the energy gap to expand as nanoparticle size decreased [46], Also, in Co_3O_4 NPs made with *cinnamon* extract, an additional absorption edge and energy gap were detected at 286.8 nm. This phenomenon is explained by the existence of substances in *cinnamon*, such as cinnamaldehyde, which can interact with metals and change the constitute of nanoparticles [47, 48].

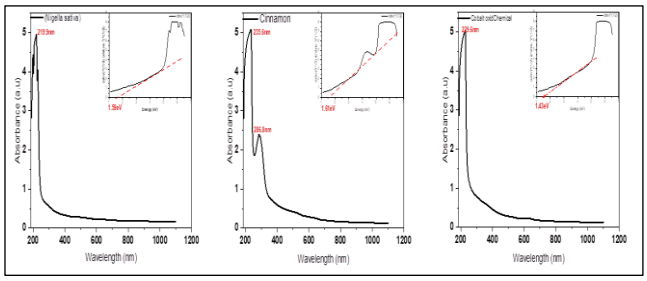


Figure 4: Ultraviolet-visible absorption spectra and the energy gap Co_3O_4 NPs with nigella sativa, cinnamon extracts, and chemical method (left to right).

FE-SEM and AFM

The general morphology of Co_3O_4 nanostructures was analyzed using FE-SEM. As shown in Fig. 5, the nanoparticles prepared by the chemical method appeared predominantly as dot-like structures with irregular edges, uniformly distributed, having a mean particle size of about 15–35 nm [49]. The FESEM images also revealed a porous structure with voids and holes of various sizes, which increases the effective surface area of the particles. Additionally, some nanoparticles formed clusters, resulting in larger aggregated particles [49, 50]. As shown in Fig. 6, the Co_3O_4 nanoparticles synthesized using *Cinnamon* extract exhibited a semi-spherical morphology, with some clustering that led to the formation of larger particles. The mean particle size for this sample ranged from 35 to 55 nm, measured using ImageJ software. Finally, as seen in Fig. 7, nanoparticles prepared with *Nigella sativa* extract displayed irregular and asymmetrical shapes, resembling uneven blocks arranged randomly. This irregular morphology is attributed to the nature of the preparation process, with a mean particle size ranging between 12.5–22.5 nm. Clusters of nanoparticles were also observed, appearing as uneven clumps.

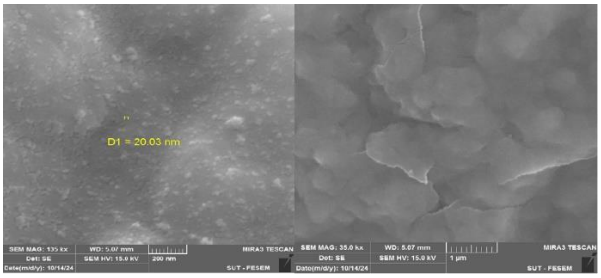


Figure 5: FE-SEM image of Co_3O_4 NPs by Chemical method.

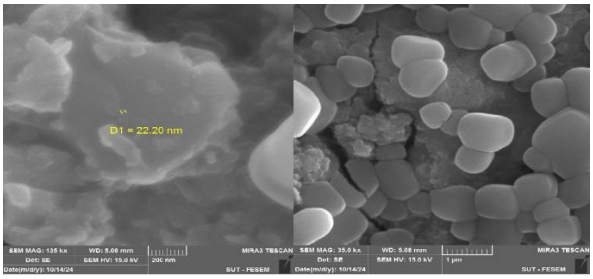


Figure (6): FE-SEM image of Co_3O_4 NPs with *Cinnamon* extract.

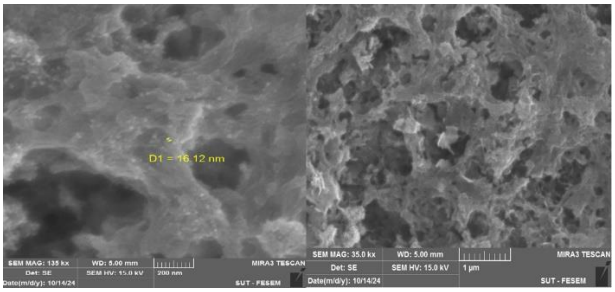


Figure (7): FE-SEM image of Co_3O_4 NPs with *Nigella sativa* extract

The AFM analysis of Co_3O_4 nanoparticles (NPs) prepared by the chemical method, as shown in Fig. 8(A), revealed a distinct nanoscale topography with a height range of up to 93.145 nm and a root mean square (RMS) roughness of 6.35 nm. The actual surface area (5.49 μm²) exceeded the projected area (4.00 μm²), reflecting an increase in effective surface area due to surface undulations. Statistical parameters, including a surface slope of 1.17 and a height distribution skewness of -0.0576, indicate slight variations with a predominance of minor depressions (51-53]. For Co_3O_4 NPs prepared using *Cinnamon* extract, as seen in Fig. 8(B), the AFM analysis exposed a nanoscale topography with a height range of up to 63.39 nm. The dark regions in the image represent pits, while the light regions represent peaks, indicating surface variations with prominent peaks and deep pits. The RMS roughness was 4.48 nm, reflecting low surface roughness, which makes the surface suitable for reducing wear and friction. The actual surface area (4.95 μm²) exceeded the projected area (4.00 μm²) due to surface roughness. Statistical parameters, such as a surface slope of 0.828 and a skewness of 0.534, suggest a predominance of low surface regions with few elevated peaks [54].In contrast, the AFM analysis of Co_3O_4 NPs prepared with *Nigella sativa* extract, as shown in Fig. 8(C), revealed a maximum height of 78.1425 nm and an RMS roughness of 3.948 nm. The actual surface area (4.695 μm²) exceeded the projected area (4.00 μm²), indicating a slight increase in the interaction area due to surface undulations. Statistical parameters, including a surface slope of 0.724 and a skewness of 0.1046, suggest a predominance of minor depressions and scattered sharp peaks. High kurtosis values indicated low to moderate surface roughness with scattered sharp features. The topographical distribution showed overall smoothness with a tilt in a specific direction (162.39°), possibly due to the manufacturing process or thermal treatment. The low roughness and positive skewness values suggest that most of the terrain consists of elevated peaks rather than deep depressions.

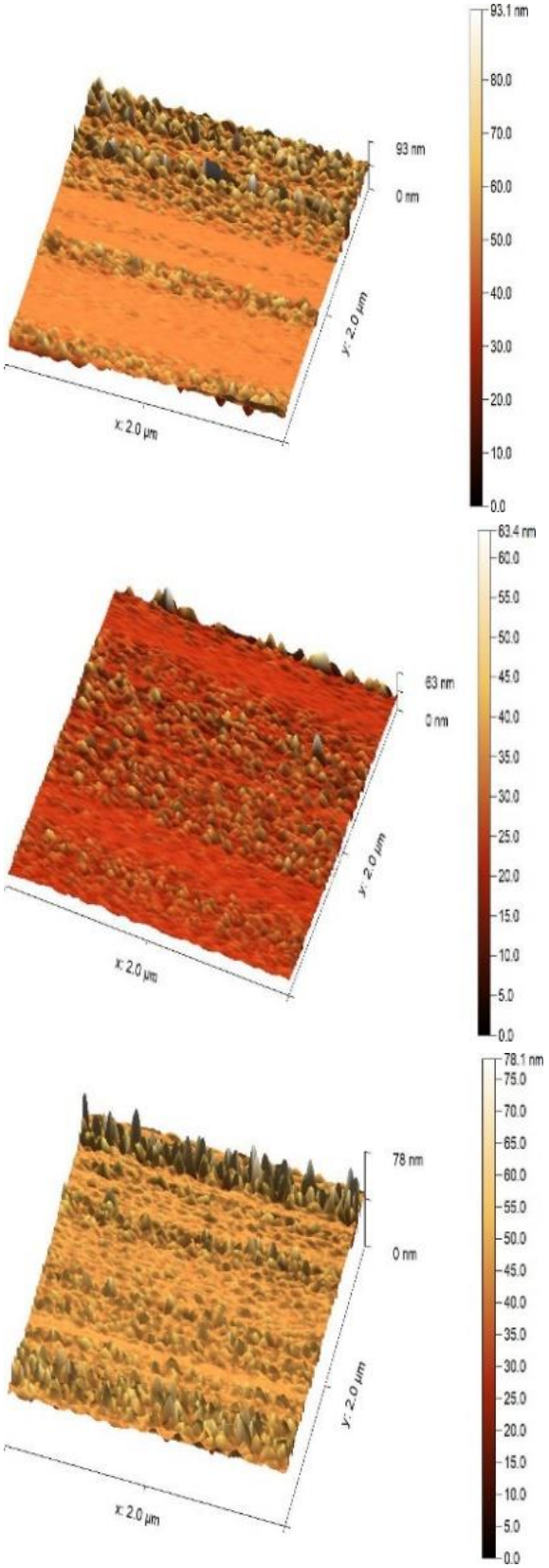


Figure (8): Atomic Force Microscopy of Co_3O_4 NPs with chemical method, *Cinnamon* extracts, and *Nigella sativa* (Left to right).

Table (2): Effect of Co_3O_4 NPs with additives on bacterial growth.

Bacteria model	*Cinnamon*	Chemical	*Nigella sativa*
Staphylococcus aureus	34	33	32
Staphylococcus epidermis	34	34	33
Escherichia coli	30	29	28
Klebsiella sp	31	34	32
Candida albicans	40	36	41

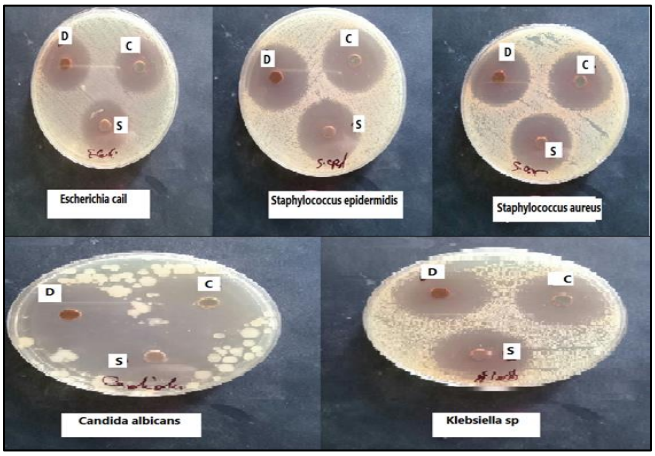


Figure (9): Antimicrobial activity of Co_3O_4 NPs synthesized by (D) *Cinnamon* extract; (C) Chemical Method: (S) *Nigella sativa*.

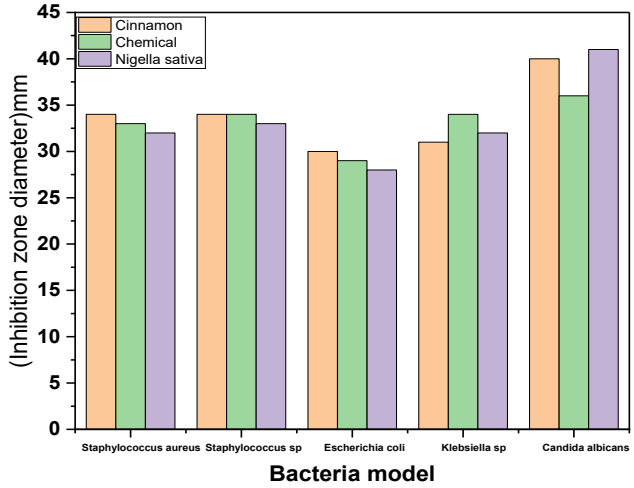


Figure (10): Effect of Co_3O_4 NPs with additives on bacterial growth.

Antibacterial Activity

Fig. 9 and Fig.10 illustrate the effects of three cobalt oxide nanoparticle (Co_3O_4 NP) samples, synthesized using plant extracts and the chemical method, on various types of bacteria and fungi. The data in Table 2, expressed as the range of the inhibition zone in millimeters surrounding the samples placed on Mueller-Hinton agar plates, reflect the antimicrobial activity of each sample. The size of the inhibition zone indicates the sample's effectiveness in inhibiting or killing microbial growth. The sample containing Co_3O_4 NPs with *Cinnamon* extract demonstrated significant effectiveness against *Staphylococcus aureus* (34 mm) and *Escherichia coli* (30 mm). This result is attributed to the presence of cinnamaldehyde, an essential oil compound with potent antibacterial properties. The synergy between cobalt oxide and cinnamaldehyde enhanced the antibacterial efficacy against Gram-positive (*S. aureus*) and Gram-negative (*E. coli*). Co_3O_4 NPs synthesized via the chemical method showed robust results, particularly against *Klebsiella sp.* (34 mm), and exhibited balanced performance against other microbes. This is likely due to the small, homogeneous particle size, which increases the active surface area and facilitates the penetration of the outer layer of Gram-negative bacteria like *Klebsiella sp.*, which are typically more resistant. Chemically synthesized nanoparticles generate a substantial amount of free oxygen radicals, such as hydroxyl radicals (•OH), superoxide anions (O_2^-), and hydrogen peroxide (H_2O_2). These radicals disrupt the outer layer of Gram-negative bacteria, damaging their DNA and proteins, ultimately leading to cell death. When Co_3O_4 NPs are combined with natural additives like cinnamaldehyde (from *Cinnamon*) or thymoquinone (from *Nigella sativa*), the production of free radicals may decrease due

to interactions between cobalt and the natural compounds. This can reduce the direct antimicrobial activity, as the concentration of active substances influences bacterial resistance. Some compounds in natural extracts may interact with cobalt oxide, partially inhibiting its direct activity. For instance, Gram-negative *Klebsiella sp.* may exhibit increased resistance to natural additives, whereas chemically synthesized Co_3O_4 NPs retain a strong ability to penetrate bacterial membranes and interact with intracellular components. *Nigella sativa*-based Co_3O_4 NPs displayed the highest effectiveness against *Candida albicans* (41 mm). This exceptional antifungal activity is attributed to thymoquinone, a potent antifungal compound that disrupts fungal cell walls, inhibits metabolic activity, and enhances the inhibitory effect through synergy with cobalt oxide. For Gram-positive bacteria like *S. aureus*, the thick peptidoglycan wall increases susceptibility to antimicrobial agents like cinnamaldehyde. Gram-negative bacteria (*Klebsiella sp.* and *E. coli*), with their additional outer membrane, exhibit greater resistance. However, Co_3O_4 NPs, especially when combined with active substances, effectively penetrate this barrier. The strong response of *Candida albicans* to *Nigella sativa* emphasizes the direct impact of thymoquinone on fungal cell walls and highlights the antifungal potency of this combination.

Conclusion

Several conclusions were drawn from the synthesis and examination of nanoparticles of cobalt oxide (Co_3O_4 NPs). Firstly, Co_3O_4 NPs were effectively created via a green synthesis technique that is both economical and ecologically benign, utilizing extracts from *Cinnamomum zeylanicum* (Cinnamon) and *Nigella sativa*. For comparison, Co_3O_4 NPs were also prepared using conventional chemical synthesis. The nanoparticles were examined with various approaches, such as using various techniques, including FTIR, XRD, FESEM, AFM, and UV-Vis spectroscopy, all of which confirmed effective fabrication of Co_3O_4 NPs. The average crystalline size and band gap energy were found to be 5.49 nm and 1.61 eV for Cinnamon-assisted NPs, 7.22 nm and 1.59 eV for *Nigella sativa*-assisted NPs, and 10.6 nm and 1.43 eV for chemically synthesized NPs. These results indicate an improvement in the electronic structure and a pronounced quantum confinement effect as particle size decreases. FTIR analysis revealed active chemical groups such as N-H, C-H, and O-H, confirming the role of plant extracts in enhancing chemical reactions and introducing functional properties to the nanoparticles. FESEM analysis demonstrated that nanoparticles synthesized with plant extracts exhibited irregular shapes and smaller sizes, leading to an increased effective surface area. This increased surface area enhanced the biological activity of the nanoparticles against bacteria and fungi. AFM analysis further confirmed distinct surface topographies for the plant-assisted NPs, including reduced surface roughness for Cinnamon-assisted NPs and increased height variations. The antimicrobial activity of Co_3O_4 NPs synthesized with natural extracts surpassed that of the chemically synthesized NPs. *Cinnamon* and *Nigella sativa* extracts exhibited significant effectiveness against both Gram-negative and Gram-positive bacteria, as well as the fungus *Candida albicans*. This superior performance is attributed to the interaction between active compounds like cinnamaldehyde (in *Cinnamon*) and thymoquinone (in *Nigella sativa*) with the nanoparticles, enhancing their membrane penetration and reactivity. In conclusion, the use of plant extracts not only facilitates the fabrication of Co_3O nanoparticles with reduced crystallite dimensions and improved energy gaps but also provides an eco-friendly and sustainable substitute for chemical techniques. This green synthesis approach holds significant potential for expanding the applications of nanoparticles in various domains, including industry, energy, and medicine.

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